

PERIODIC TABLE OF THE ELEMENTS

1	H 1.0079											2	He 4.0026									
3	4											9	10									
Li 6.941	Be 9.012											F 19.00	Ne 20.179									
11	12											17	18									
Na 22.99	Mg 24.30											Cl 35.453	Ar 39.948									
19	20	21	22	23	24	25	26	27	28	29	30	31	32	33	34	35	36					
K 39.10	Ca 40.08	Sc 44.96	Ti 47.90	V 50.94	Cr 52.00	Mn 54.938	Fe 55.85	Co 58.93	Ni 58.69	Cu 63.55	Zn 65.39	Ga 69.72	Ge 72.59	As 74.92	Se 78.96	Br 79.90	Kr 83.80					
37	38	39	40	41	42	43	44	45	46	47	48	49	50	51	52	53	54					
Rb 85.47	Sr 87.62	Y 88.91	Zr 91.22	Nb 92.91	Mo 95.94	Tc (98)	Ru 101.1	Rh 102.91	Pd 106.42	Ag 107.87	Cd 112.41	In 114.82	Sn 118.71	Sb 121.75	Te 127.60	I 126.91	Xe 131.29					
55	56	57	72	73	74	75	76	77	78	79	80	81	82	83	84	85	86					
Cs 132.91	Ba 137.33	*La 138.91	Hf 178.49	Ta 180.95	W 183.85	Re 186.21	Os 190.2	Ir 192.2	Pt 195.08	Au 196.97	Hg 200.59	Tl 204.38	Pb 207.2	Bi 208.98	Po (209)	At (210)	Rn (222)					
87	88	89	104	105	106	107	108	109	110	111	112	§ Not yet named										
Fr (223)	Ra 226.02	†Ac 227.03	Rf (261)	Db (262)	Sg (266)	Bh (264)	Hs (277)	Mt (268)	Ds (271)	Rg (272)	S (277)											
													67	68	69	70	71					
													Ho 164.93	Er 167.26	Tm 168.93	Yb 173.04	Lu 174.97					
													99	100	101	102	103					
													Es (252)	Fm (257)	Md (258)	No (259)	Lr (262)					
													66	65	64	63	62	61	60	59	58	
													Dy 162.50	Tb 158.93	Gd 157.25	Eu 151.97	Sm 150.4	Pm (145)	Nd 144.24	Pr 140.91	Ce 140.12	
													98	97	96	95	94	93	92	91	90	
													Cf (251)	Bk (247)	Cm (247)	Am (243)	Pu (244)	Np 237.05	U 238.03	Pa 231.04	Th 232.04	
													*Lanthanide Series					†Actinide Series				

Practice SAT Chemistry Trial

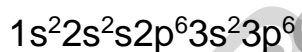
1. What is an excited state electrons configuration of fluoride ion?
 - A. $1s^2 2s^2 2p^6 3s^2$
 - B. $1s^2 2s^2 2p^6 3s^2 3p^5$
 - C. $1s^2 2s^2 2p^5$
 - D. $1s^2 2s^2 2p^5 3s^1$
 - E. $1s^2 2s^2 2p^4 3s^1$

2. The excited electron of sodium atom locates in orbital
 - A. 1s
 - B. 2s
 - C. 2p
 - D. 3s
 - E. 3p

3. A rectangular block of aluminum alloy weighs 20.45 g. The dimensions of the block are 8.4 cm by 5.5 cm by 4.6 cm. From this data, what is the density of this alloy?
 - A. 0.0962 g/cm^3
 - B. 0.0960 g/cm^3
 - C. 0.096 g/cm^3
 - D. 0.09 g/cm^3
 - E. 0.9 g/cm^3

4. What is the number of moles in 15.0 g CO_2 ?
(Atomic mass of C = 12.011, O = 16.00 g/mol)
 - A. 0.341
 - B. 0.34
 - C. 0.3
 - D. 0.340
 - E. 0.35

5. What statement is/are true for $1s^2 2s^2 2p^5 3s^1$
- I. Excited state electrons configuration of neon atom
 - II. Excited state electrons configuration of magnesium ion
 - III. Excited state electrons configuration of oxygen atom
- A. I
 - B. II
 - C. I and II
 - D. I and III
 - E. I, II and III



6. The electrons configuration above represents following EXCEPT
- A. Ground state electrons configuration of sulfide ion
 - B. Ground state electrons configuration of argon atom
 - C. Ground state electrons configuration of calcium ion
 - D. Ground state electrons configuration of potassium ion
 - E. Ground state electrons configuration of potassium atom
7. What is the mass of in gram of 0.031 mol NO_2 ?
(Atomic mass of N = 14.00, O = 16.00 g/mol)
- A. 1.426
 - B. 1.43
 - C. 1.4
 - D. 1.40
 - E. 1.42